



Course Syllabus
According to JORDAN National Qualification
Framework (JNQF)

Course Name: Physical Chemistry laboratory (2)

Course Number: 11013144

General Course Information:

Course title	Physical Chemistry laboratory (2)
Course number	11013144
Credit hours	2 Credit hours
Education type	Face-to-Face
Prerequisites/corequisites	11012242+11013143 or synchronous
Academic Program	Bachelor
Program code	01
Faculty	Faculty of science
Department	Department of Chemistry
Level of course	2nd year
Academic year /semester	First semester 2022/2023
Awarded qualification	BSc degree of chemistry
Other department(s) involved in teaching the course	None
Language of instruction	English
Date of production/revision	2022/2023

Course Coordinator:

Coordinator's name	Ms. Daren Hemedat
Office No	4228
Office Phone extension number	2635
Office Hours	
Email	darenhmedat@yahoo.com

Other Instructors:

Instructor name	
Office No	
Office Phone extension number	
Office Hours	
Email	

Course Description (English/Arabic):

English	Experiments in electrolytic solutions (determination of solubility product constant of silver chloride salt and dissociation constant of weak acid by measuring conductivity); experiments in kinetic chemistry (determination the order of chemical reaction and rate law for several types of chemical reactions based on different methods); experiments in surface chemistry (Adsorption).
Arabic	تجارب في المحاليل الكهروlyتية (تحديد ثابت الذوبان لملاح كلوريد الفضة و ثابت التفكك لحمض ضعيف عن طريق قياس الموصلية); تجارب في الكيمياء الحركية (تحديد درجة التفاعل الكيميائي و قانون معدل سرعة التفاعل الكيميائي لعدد من التفاعلات الكيميائية باستخدام طرق مختلفة); تجارب في كيمياء السطح (الادمصاص).

Textbook: Author(s), Title, Publisher, Edition, Year, Book website.

1. Practical physical Chemistry (2) Laboratory manual, Isra University 2018-2019

References: Author(s), Title, Publisher, Edition, Year, Book website.

Required book (s), assigned reading and audio-visuals:

1. Collected Manual, Department of Applied Chemical Science

Course Educational Objectives (CEOs):

1.	To reinforce the material the student have learned the physical chemistry 2 class (11013143) and to give the students the chance to apply what they have learned practically
2.	The students will learn some new experimental techniques that are necessary for them to become an effective chemists and researchers.
3.	The student will learn how to write a lab report in a professional manner.
4.	Able to design work both individually and as part of team

Intended Learning Outcomes (ILO's):

1.	Subject Intended learning outcomes (ILOs) describe what students are expected to know and be able to do at the end of the course. These outcomes are related to the knowledge, skill and competence that students acquire:	Relationship to CEOs	Contribution to PLOs	Bloom Taxonomy Levels*	Descriptors**
2. A	Knowledge and Understanding:				
3. A1	Student will be able to recognize fundamentals of physical chemistry including chemical kinetics, collide chemistry spectroscopy and electrochemistry	1	1	1	1
4. B	Intellectual skills:				
5. B1	Student will be able to perform calculations related to experiments this may include sketching graphs and solving algebraic equations.	3	2	4	s
6. C	Subject specific skills:				
7. C1	Student will be able to applied procedures in chemical reactions to determine (the rate and the order of reaction , equilibrium distribution of organic acid between an aqueous	2,3	2,3	3,4	S,c

	solution, dissociation constant of weak electrolyte based on conductometric method and specific area of charcoal) and applying adsorption method .				
8. D	Transferable skills:				
9. D1	Student will be learn how to work individually and with partners effectively	4	7	3	c

*Bloom Taxonomy Levels

Level #	1	2	3	4	5	6
Level Name	Knowledge	Comprehension	Application	Analysis	Evaluation	Synthesis

** Descriptor (National Qualification Framework Descriptors): K : Knowledge, S: Skill, C: Competency.

Program Learning Outcome (PLOs):

Program Learning Outcomes describe what students are expected to know and be able to do by the time of graduation. These relate to the knowledge, skills, and behaviours that students acquire as they progress through the program. A graduate of the (Chemistry) program will demonstrate:	Descriptors**		
	K	S	C

1.	An ability to acquire and apply new knowledge as required across different fields of chemistry, using appropriate learning strategies.	✓		
2.	An ability to identify, formulate, and solve problems by applying principles and theories of chemistry, science and mathematics based on critical thinking.		✓	
3.	An ability to develop and conduct appropriate experimentation, analyze, interpret data, and draw conclusions.		✓	
4.	An ability to apply scientific principles and theories of chemistry to serve community in health, economic and environmental sectors.			✓
5.	An ability to communicate effectively with a wide range of audiences			✓
6.	An ability to recognize ethical and professional responsibilities in the field of chemistry, and make informed judgments that consider the impact of chemistry in global, economic, environmental and societal contexts.			✓
7.	An ability to function effectively as a part of a team, take on leadership positions, create a collaborative and inclusive environment, establish goals, plan tasks, and meet objectives			✓

**** Descriptors according to the national qualifications framework (K: knowledge, S: skill, C: Competency)**

Weekly Schedule (please choose the type of teaching)

✓ (4 hrs Face – To - Face)

Hybrid (2 Lectures Face – To - Face +1 Lecture Asynchronous)

Hybrid (1 Lectures Face – To - Face +1 Lecture Asynchronous)

Online (2 Lectures Synchronous +1 lecture Asynchronous)

Week	four Hours (face to face)	Second Hour (.....)	Third Hour (.....)	Ach. ILOs	Ach. PLOs	Descriptors*
Week 1	Laboratory safety rules			C1	2,3	s
Week 2	Determination of equilibrium of adsorption of an organic acid by activated carbon in aqueous medium			C1 & B1 & D1	1,2,3,7,	s
Week 3	Kinetics of the catalyzed Decomposition of Hydrogen Peroxide			A1 & B1 & D1	1,2,3,7	s
Week 4	Effect of ionic strength on solubility of benzoic acid in water.			C1 & B1 & D1	1,2,3,7,	s
Week 5	Effect of ionic strength on rate of reaction			C1 & B1 & D1	1,2,3,7,	s
Week 6	Reaction of ethyl			C1 & B1 & D1	1,2,3,7,	s

	acetate with hydroxyl ion followed by electrical conductance			D1		
Week 7	Dissociation Constants and Solubility from Conductance of Electrolyte Solutions					
Week 8	Midterm exam			C1 & B1 & D1	1,2,3 7,	s
Week 9	Conductance behaviour of strong and weak electrolytes.			C1 & B1 & D1	1,2,3 7,	s
Week10	Determination of The Solubility Product of A Sparingly Soluble Salt			C1 & B1 & D1	1,2,3 7,	s
Week11	Acid dissociation constant of methyl red.			C1 & B1 & A1 & D1	1,2,3 7,	s
Week12	Chemical Kinetic: Hydrolysis of Sucrose			C1 & B1 & D1	1,2,3 7,	s
Week13	Final exam					

* K: Knowledge, S: Skills, C: Competency

Teaching Methods and Assignments:

Development of ILOs is promoted through the following teaching and learning methods:

- (4 hrs Face – To - Face)
- Practice Labs
- Discussion Forums
- Quizzes
- Other Interactive online activities
- Reports

Course Policies:

A- Attendance policies:

The maximum allowed absences is 15% of the lectures.

B- Absences from exams and handing in assignments on time:

Midterm exam can be retaken based on approval of excuse by the instructor's discretion.

Not handing assignment on time will incur penalties.

C- Academic Health and safety procedures

D- Honesty policy regarding cheating, plagiarism, and misbehaviour:

Cheating, plagiarism, misbehaviour will result in zero grade and further disciplinary actions may be taken.

E- Grading policy:

- All homework is to be posted online through the e-learning system.
- Exams will be marked within 72 hours and the marked exam papers will be handed to the students.
- Online Activities (Course Videos, Practice labs, Discussion Forums, Quizzes) **20%**
- Midterm **30 %**
- Final Exam **50 %**

F- Available university services that support achievement in the course: **E-Learning Platform, Labs, Library.**

Required equipment:

- **PC / Laptop with webcam and mic**
- **Internet Connection**
- **Access to the IU E-Learning Platform at: <https://elearn.iu.edu.jo/>**
- **E-learning plan**
- Satisfaction questionnaires for online and face-to-face learning
- Software for e-learning
- Training

Assessment Tools implemented in the course:

- Final Exam
- Midterm Exam
- Quizzes
- Homework
- Practice Labs
- Discussion Forums
- Periodic reports for learning assessment
- Improvement plans for online or face-to-face teaching

Responsible Persons and their Signatures:

Course Coordinator	M.Sc. Dareen Hmedat	Completed Date	16/10/2022
		Signature	<i>Dareen Hmedat</i>
Received by (Department Head)	Dr. Manal Al Khabas	Received Date	16/10/2022
		Signature	<i>Manal Al Khabas</i>